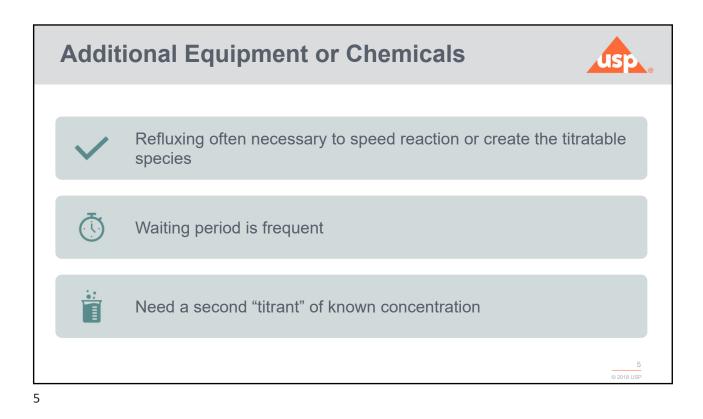
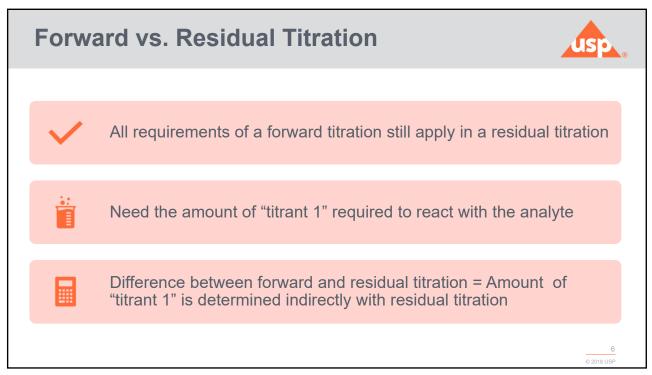
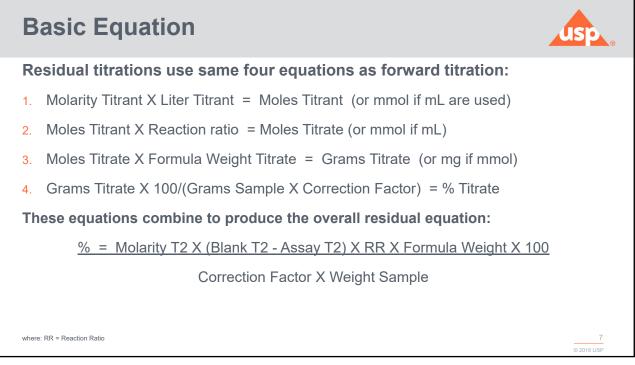
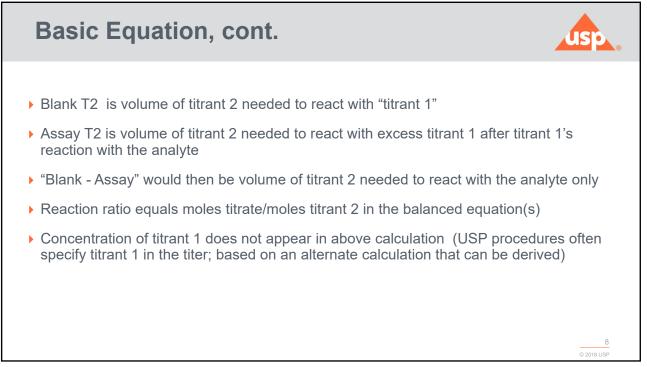


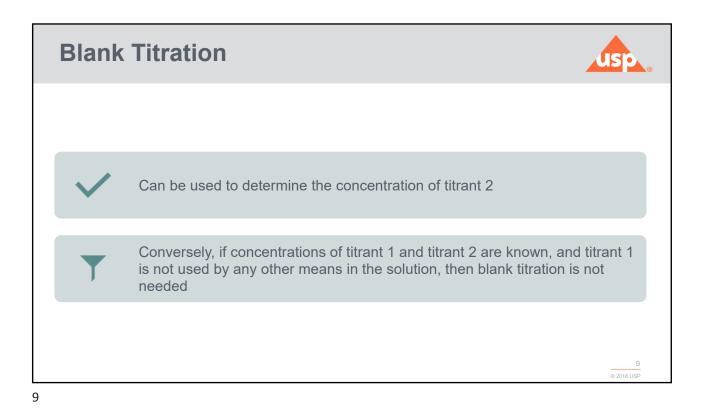
The Reasons		
	Used to determine an analyte where the forward reaction is slow or thermodynamically unfavorable (Hydrolysis of esters)	
Ĩ	Used to determine an analyte when the analyte does not react directly with the titrant (Bromination reactions)	
Â	Solubility problems	
~	Secondary reactions	
898	Endpoint determination not possible (Metal ion titrations)	
×	Several reasons may apply	
	4 © 2018 USP	

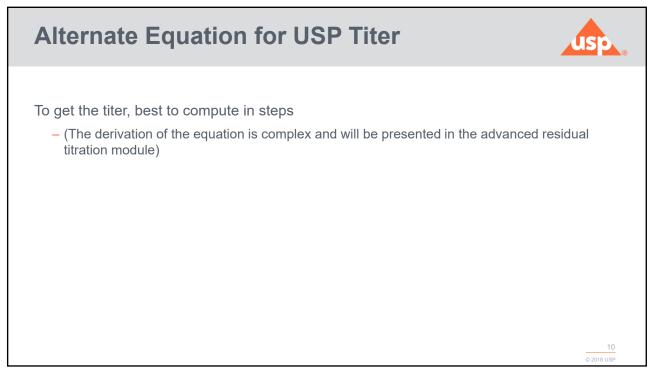


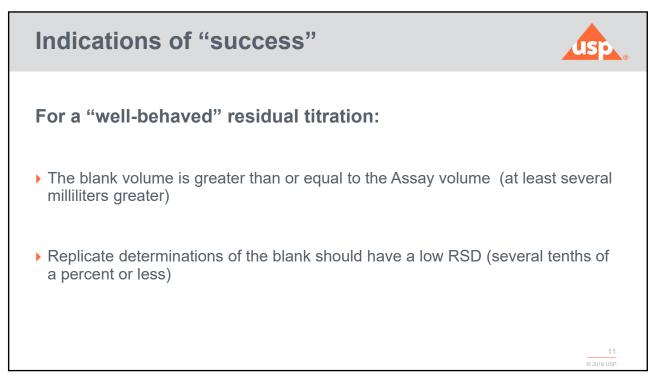


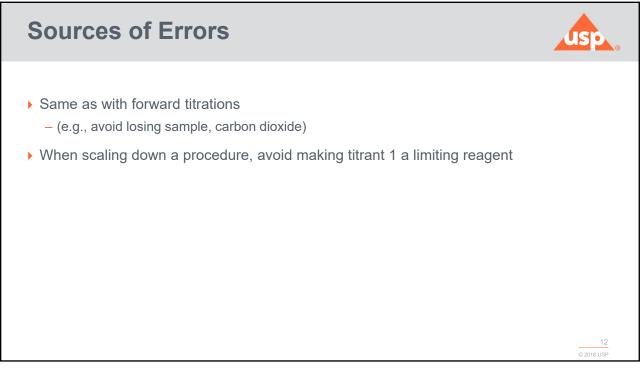


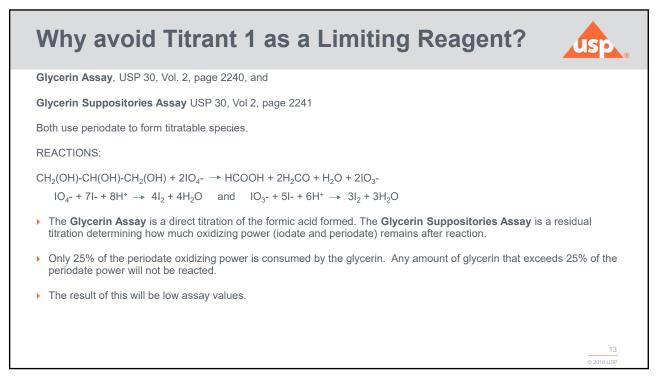


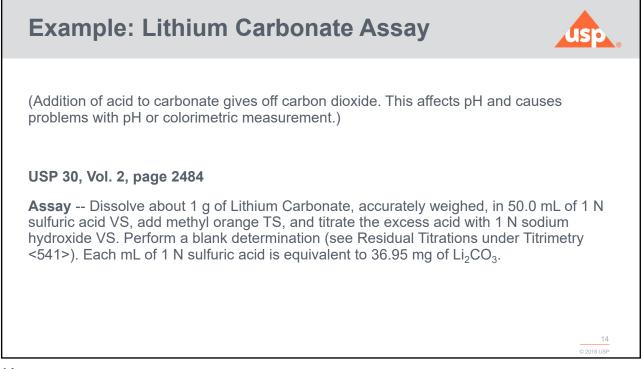




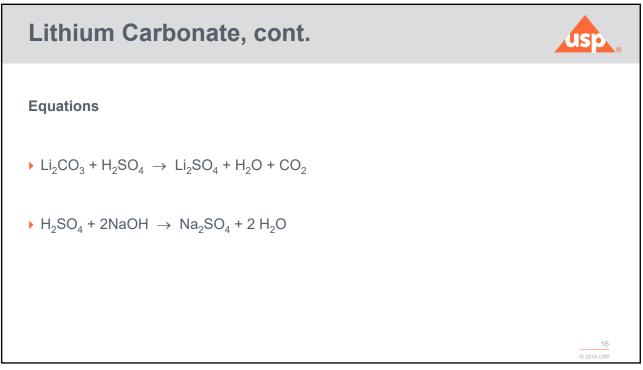


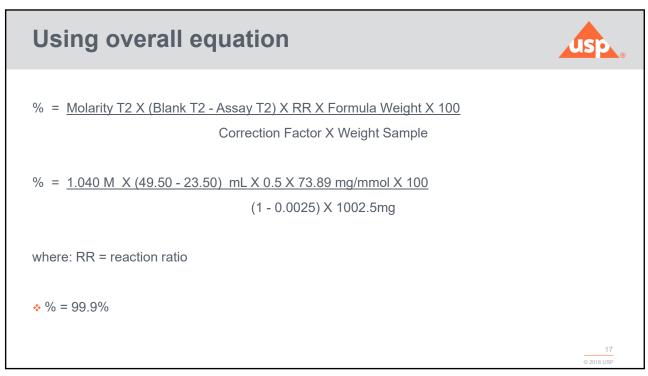


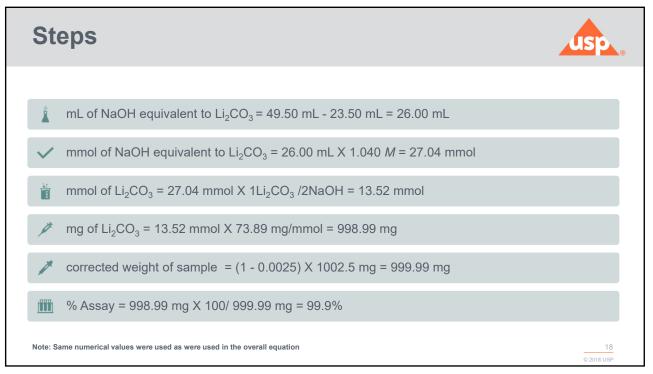




Lithium Carbonate, cont. Data Weight Lithium Carbonate: 1.0025 g LOD: 0.25% Molarity sodium hydroxide VS: 1.040 M (N) Volume of sodium hydroxide in Assay: 23.50 mL Volume of sodium hydroxide in Blank: 49.50 mL T1 = sulfuric acid, T2 = sodium hydroxide







Conclusion

Residual Titrations look difficult, but they are not.

If you:

- know the chemistry (including balanced reactions)
- know what to expect, and
- know the pitfalls and the sources of error

you should succeed.

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